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## 1 through 5

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SnCO tin(II) carbonate  
sodium hydrogen carbonate NaHCO 1. 2. 3. 4. 5. 6. 8. 9. 10. 11.

vanadium(V) oxide H<sub>2</sub>O  
 dihydrogen monoxide  
 (NH<sub>4</sub>)<sub>2</sub>CO ammonium  
 oxalate 4 224  
 polonium(VI) thiocyanate  
 Po(SCN)<sub>6</sub> tetraphosphorus  
 decaoxide P<sub>4</sub>O<sub>10</sub> 12.  
 13. 15. 16. 17. 18. 19.  
 manganese(VII) oxide  
 copper(II) dihydrogen  
 phosphate

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 following single  
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 Smith question requiring  
 the determination of  
 products of chemical  
 reactions (the Equations  
 section) five of these  
 eight equations to answer  
 (Note: Additional  
 responses will not be  
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Note: Not all of the reactions will occur. For those that do not, write no reac-

tion. l. A piece of copper is dropped into a container of water. No reaction 2. Liquid bromine is added to a container of sodium iodide crystals.  $\text{Br (l)} + 2\text{NaI(s)} \rightarrow 2\text{NaBr(s)} + \text{I}_2\text{(s)}$  3. An aluminum strip is immersed in a solution of silver nitrate.

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the determination of products of chemical reactions (the Equations section) five of these eight equations to answer (Note: Additional responses will not be graded or counted If all [PDF] Ultimate Chemical Equations Handbook Answer Key

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